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| --- | --- |
| ASC-ColPos-Horizontal small | **Year 12 Chemistry****Test #1 (Atomic Structure & Bonding)****Weighting: 2% Time: 50 minutes** |

Name: **ANSWERS** Mark = \_\_\_\_\_ / 46

**Part 1: Multiple Choice Section 9 marks**

*Answer by placing a cross through, or a circle around, the letter of the most correct answer.*

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1. The full symbol for a particular ion of arsenic-71 is .

 Which of the following best describes the composition of this ion?

 **A. 33 protons, 38 neutrons and 30 electrons.**

 B. 33 protons, 38 neutrons and 33 electrons.

 C. 33 protons, 71 neutrons and 30 electrons.

 D. 38 protons, 33 neutrons and 38 electrons.

2. Consider the Lewis structure for a polyatomic anion of element X:



 Element X is likely to be in:

 **A. group 14 24 valence electrons in diagram**

 B. group 15 **18 from oxygen (6 each)**

 C. group 16 **2 from elsewhere (­-2 charge)**

 D. group 17 **∴ 4 from X, ∴ group 14**

3. Which of the following combinations of elements are listed in **increasing** order of electronegativity?

 A. phosphorous, nitrogen, oxygen, sodium, magnesium

 B. sodium, phosphorous, oxygen, nitrogen, fluorine **🡩 across periods**

 **C.**  **sodium, phosphorous, nitrogen, oxygen, fluorine 🡫 down groups**

 D. sodium, magnesium, oxygen, nitrogen, fluorine

4. Which of the following correctly identifies the shapes of the following molecules?

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
|  | SiH4 | PH3 | H2S | HC |
|  A. | planar | pyramidal | linear | linear |
|  B. | tetrahedral | trigonal planar | linear | bent |
|  C. | pyramidal | trigonal planar | bent | linear |
|  **D.**  | **tetrahedral** | **pyramidal** | **bent** | **linear** |

5. Which of the substances listed below have both a trigonal planar shape and a dipole?

 I NH3 **pyramidal**

 II H2CO **trig planar & polar**

 III SO3 **trig planar & non-polar**

 IV PC3 **pyramidal**

 V COC2 **trig planar & polar**

 A. I and IV only

 B. II and III only

 **C.**  **II and V only**

 D. II, III and V only

6. The molecules hydrogen sulfide (H2S), methanamine (CH3NH2), and oxygen (O2),

 have similar molar mass. Which of the following lists the gases in **ascending** order of

 boiling point?

 A. H2S, O2, CH3NH2 **For molecules of similar size, strength of IMF is:**

 **B.**  **O2, H2S, CH3NH2 dispersion < dipole-dipole < hydrogen bonding**

 C. O2, CH3NH2, H2S

 D. CH3NH2, H2S, O2

7. Which of the following statements is **incorrect** for an ionic substance?

 A. The substance will have a high melting point because of the strong electrostatic

 attraction between oppositely charged ions.

 B. When heated sufficiently charged particles can move and allow the passage of an

 electric current through the substance.

 **C. When dissolved in water the ionic lattice breaks up and makes electrons**

 **available to allow the passage of an electric current through the solution.**

 D. When the ions in the lattice are forced to move, electrostatic repulsion tends to

 make the solid shatter.

 **It is the cations and anions that are the charge carriers in aqueous ionic solution**

8. A covalent bond is best described as:

 A. a bond between two non-metallic elements.

 B. the sharing of electrons between two atoms.

 **C. the attraction between the nuclei of adjacent atoms and their shared electrons.**

 D. either a polar or non-polar bond.

9. The table below gives four consecutive ionisation energies (in MJ mol­–1) for an element in

 the third period of the Periodic Table.

|  |  |  |  |
| --- | --- | --- | --- |
| 1st | 2nd | 3rd | 4th |
| 0.425 | 3.058 | 4.418 | 5.883 |

 **Significant jump, ∴ one valence electron**

 **Third period, ∴ three electron shells**

 The ground state electron configuration for atoms of this element is:

 A. 2,1

 **B.**  **2,8,1**

 C. 2,8,4

 D. 2,8,8,1

**End of Part One**

**Part Two: Short Answer Section 37 marks**

*Write all answers in the spaces provided.*

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**Question 10 (5 marks)**

Consider the elements chlorine, magnesium, neon and phosphorus.

(a) Rank the elements in **increasing** order of first ionisation energy.

 **Mg < P < C < Ne ✓**

(1 mark)

(b) Explain your reasoning for your ranking in (a) above.

 *Note that simply stating a trend is not an explanation.*

 **Across (left to right) a period, first ionisation energy increases ✓**

 **because valence electrons in the same shell are being**

 **attracted to a more positive nucleus. ✓**

 **∴ Mg < P < C (< Ar)**

 **Down a group, first ionisation decreases ✓**

 **because valence electrons are in a new shell**

 **at a greater distance from the nucleus ✓**

 **∴ C (< Ar) < Ne**

(4 marks)

**Question 11 (4 marks)**

|  |  |
| --- | --- |
| The Lewis structure for hydrogen peroxide, H2O2, is shown to the right. The O-O bond can be described as non-polar whereas the O-H bond can be described as polar. Explain the meaning of the terms polar and non-polar in relation to covalent bonds. In your response explain why there is a difference in polarity of the O-O and O-H bonds. |  |

**A non-polar bond is one in which there is an even distribution of charge. ✓**

**For example, in the O-O bond, the electrons are shared between two atoms of**

**equal electronegativity, so they are evenly distributed between them. ✓**

**A polar bond is one in which there is an uneven distribution of charge. ✓**

**For example, in the O-H bond, the electrons are more strongly attracted to**

**the O, which has greater electronegativity than H. The O, therefore, acquires a**

**partial negative charge (δ­-) and H acquires a partial positive charge (δ+). ✓**

**Question 12 (9 marks)**

Complete the table given below by:

- drawing Lewis structures, representing all valence shell electron pairs as : or as –

- naming or drawing the molecular shapes

- identifying the molecules as either polar or non-polar

|  |  |  |  |
| --- | --- | --- | --- |
| *Formula* | *Lewis structure* | *Molecular shape* | *Polar/non-polar* |
| CH2C2 |  | **tetrahedral** | **polar** |
| PBr3 |  | **trigonal pyramidal****(or pyramidal)** | **polar** |
| N2O(NNO) | Picture 2**or**Picture 3 | **linear** | **polar** |

**✓ each**

**Question 13 (3 marks)**

Consider the following information:

- element X is a silvery-grey solid at room temperature. It melts at 660°C and is a good thermal

 and electrical conductor

- element Y is a red liquid at room temperature and a non-conductor of electricity in any state

- element X and Y combine to form a compound that has a low melting point (98°C) and, when

 molten, is a non-conductor of electricity

(a) Classify the compound of X and Y as ionic, metallic, molecular or covalent network.

 **Molecular ✓**

(1 mark)

(b) Provide supporting reasons for your classification in (a) above.

 **Non-conductor of electricity when molten, so particles are not charged**

 **∴ cannot be metallic or ionic. ✓**

 **Low melting point, so particles not bound by strong bonds/forces**

 **∴ cannot be covalent network. ✓**

 **By elimination, must be covalent molecular.**

(2 marks)

**Question 14 (4 marks)**

(a) Explain what is meant by the term ‘hydrogen bond’.

 **A hydrogen bond is an intermolecular force of attraction between a non-bonding**

 **pair of electrons on a O/N/F in one molecule and a partially positive hydrogen in an**

 **adjacent molecule (where it is covalently bonded to O/N/F) ✓✓**

(2 marks)

(b) Consider the substances listed below. Draw a circle around those that could possibly

 form a hydrogen bond with water.

|  |  |  |
| --- | --- | --- |
| **hydrogen fluoride** | phosphine (PH3) | **ammonia** |
| trichloromethane (CHC3)  | **methanal (CH2O)** |

 **2 correct ✓, 3 correct ✓✓** (2 marks)

**Question 15 (6 marks)**

Consider the following substances and their melting points.

HBr – 86⁰C Br2 5⁰C CBr4 90⁰C

Explain the difference in melting points for the each of the following pairs of substances:

(a) Br2 and CBr4

 **Both are non-polar molecules with only dispersion forces between molecules. ✓**

 **Dispersion forces increase in strength with an increase in the number of**

 **protons and electrons in the molecule. ✓**

 **CBr4 has a greater number of electrons than Br2 and so its dispersion forces**

 **are stronger. Greater energy is required to overcome these stronger forces**

 **of attraction and so CBr4 has a higher boiling point. ✓**

 (3 marks)

(b) HBr and Br2

 **Br2 is a non-polar molecule with only dispersion forces between molecules. **

 **HBr is a polar molecule with dispersion and dipole-dipole forces between molecules. **

 **However, Br2 has a greater number of electrons than HBr**

 **so its dispersion forces are stronger – in this case stronger than the**

 **dispersion and dipole-dipole forces between HBr molecules. **

 **Therefore, more energy needs to be supplied to overcome IMF between**

 **molecules of Br2 than HBr.**

(3 marks)

**Question 16 (6 marks)**

Consider the hydrides of the group 15 elements.

(a) On the axes below sketch a qualitative graph for the boiling points of the hydrides in this

 group.

Boiling point

(°C)

 **Note: graph need only be ‘qualitative’**

 **and show PH3 as lowest of the series**

 **✓ NH3 > PH3,**

 **✓ BiH3 > SbH3 > AsH3 > PH3**

 NH3 PH3 AsH3 SbH3 BiH3

 (2 marks)

(b) Account for the shape of your graph.

 **NH3 > PH3**

 **N atom has high electronegativity so N-H bond is highly polar. **

 **Hydrogen bonds form between NH3 molecules, so, despite its**

 **small size, NH3 has a higher boiling point than PH3. **

 **BiH3 > SbH3 > AsH3 > PH3**

 **Molecular polarity decreases, so dipole-dipole forces become weaker. ✓**

 **However, the number of electrons per molecule increases, which increases**

 **the strength of dispersion forces and so boiling points increases. **

(4 marks)

**End of Test**